

## *Oxidation Number Practice Worksheet Answers*

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### Oxidation Number Practice Worksheet Answers

Rules for Assigning Oxidation Numbers 1. The oxidation number of any uncombined element is 0. 2. The oxidation number of a monatomic ion equals the charge on the ion. 3. The more-electronegative element in a binary compound is assigned the number equal to the charge it would have if it were an ion. 4.

### Oxidation Numbers Worksheet - brookville.k12.oh.us

Practice Problems: Redox Reactions (Answer Key) Determine the oxidation number of the elements in each of the following compounds: a.  $\text{H}_2\text{CO}_3$  H: +1, O: -2, C: +4

### Practice Problems: Redox Reactions (Answer Key)

Worksheet 25 - Oxidation/Reduction Reactions Oxidation number rules: Elements have an oxidation number of 0 Group I and II - In addition to the elemental oxidation state of 0, Group I has an oxidation state of +1 and Group II has an oxidation state of +2. Hydrogen -usually +1, except when bonded to Group I or Group II, when it forms hydrides, -1.

### Worksheet 25 - Oxidation/Reduction Reactions 0 II +1 +2 -2 -1

There are two oxygens, and oxygen has an oxidation number of -2, according to rule 3. Therefore, sulfur should have an oxidation number of +4, because  $+4 + (2 * (-2)) = 0$ . For the following quiz, please read each question carefully. Use the above summary and example to help you determine the answer.

### Oxidation Numbers Quiz - Free Math worksheets, Free ...

Showing top 8 worksheets in the category - Oxidation. Some of the worksheets displayed are Work 7, Oxidation state work and answers, Oxidation number exercise, Work 1 determination of oxidation number or valence, Work oxidation numbers name, Oxidation number work 1 write the rule next to your, Chapter 20 work redox, Key review work on balancing redox equations.

### Oxidation Worksheets - Printable Worksheets

Exercises - Give the oxidation number for the following atoms: NaF Na = +1 IF<sub>3</sub> I = +3 ClF<sub>2</sub> Cl = +1 SF<sub>4</sub> S = +4 PF<sub>3</sub> P = +3 SF<sub>6</sub> S = +4 PF<sub>5</sub> P = +5 PF<sub>6</sub> P = +3 W<sub>2</sub>F<sub>9</sub> W = +3 OF<sub>2</sub> O = +2 NF<sub>3</sub> N = +3 F<sub>2</sub> F = 0 Rule 3 The metal of groups IA have an oxidation number of +1 The metal of groups IIA have an oxidation number of +2 Sc, Y and Al have an oxidation number of +3.

### Oxidation Number Exercise - Roane State Community College

This quiz and worksheet will help you check your understanding of oxidation numbers and how to assign them to atoms in molecules. You will be asked to assign oxidation numbers to elements in a ...

### Quiz & Worksheet - How to Assign Oxidation Numbers to ...

Oxidation Numbers. Showing top 8 worksheets in the category - Oxidation Numbers. Some of the worksheets displayed are Work oxidation numbers name, Work 25, Work assigning oxidation numbers, Chapter 20 work redox, Oxidation number exercise, Oxidation reduction handout, Work 1 determination of oxidation number or valence, Work 7.

### Oxidation Numbers Worksheets - Printable Worksheets

OXIDATION NUMBER EXERCISE This is an exercise in determining the oxidation numbers in ions and compounds. Calculate the oxidation numbers of all the elements using the rules discussed in class. Mouse over the formula to reveal the answers. NOTE: This page works ONLY in Internet Explorer and not Netscape Navigator.

### OXIDATION NUMBER EXERCISE - College of DuPage

oxidation numbers of element X are A. +1 and +2 B. +2 and +3 C. +1 and +3 D. +2 and +4 8. Oxygen has a positive oxidation number in the compound A. H<sub>2</sub>O B. H<sub>2</sub>O<sub>2</sub> C. OF<sub>2</sub> D. IO<sub>2</sub> 9. What is the oxidation number of sulfur in H<sub>2</sub>SO<sub>4</sub>? A. 0 B. 2 C. +6 D. +4 10. In the equation  $\text{Cu(s)} + 2\text{Ag}^+(\text{aq})$

$\text{Cu}^{2+}(\text{aq}) + 2\text{Ag}(\text{s})$ , the oxidizing agent is A.  $\text{Cu}^0$  B.  $\text{Ag}^+$  C.  $\text{Cu}^{2+}$  D.  $\text{Ag}^0$  11.

### Redox practice worksheet - Imghs.org

Thus the oxidation number of Cl in the  $\text{Cl}^-$  ion is -1, that for Mg in the  $\text{Mg}^{2+}$  ion is +2, and that for oxygen in  $\text{O}^{2-}$  ion is -2. • The sum of the oxidation numbers in a compound is zero if neutral, or equal to the charge if an ion. • The oxidation number of alkali metals in compounds is +1, and that of alkaline earths in compounds is +2. The oxidation number of F is -1 in all its compounds. • The oxidation number of H is +1 in most compounds.

### Redox Balancing Worksheet - Strongsville City Schools

1. Give the oxidation numbers of all the elements in the following molecules and ions: a.  $\text{SO}$ ,  $\text{SO}_2$ ,  $\text{SO}_3$ ,  $\text{SO}_3^{2-}$ ,  $\text{SO}_4^{2-}$  b.  $\text{ClO}_2$ ,  $\text{ClO}^-$ ,  $\text{ClO}_2^-$ ,  $\text{ClO}_3^-$ ,  $\text{ClO}_4^-$  c.  $\text{N}_2\text{O}$ ,  $\text{NO}$ ,  $\text{NO}_2$ ,  $\text{N}_2\text{O}_4$ ,  $\text{N}_2\text{O}_5$ ,  $\text{NO}_2^-$ ,  $\text{NO}_3^-$  2. Determine the oxidation number of the sulfur atom: a.  $\text{H}_2\text{S}$  b.  $\text{S}$  c.  $\text{H}_2\text{SO}_4$  d.  $\text{S}^{2-}$  e.  $\text{HS}^-$  f.  $\text{SO}_2$  g.  $\text{SO}_3$  3.

### Worksheet: Oxidation Numbers Name

Practice Problems: Redox Reactions. Determine the oxidation number of the elements in each of the following compounds: a.  $\text{H}_2\text{CO}_3$  b.  $\text{N}_2$  c.  $\text{Zn}(\text{OH})_2$  d.  $\text{NO}_2$  e.  $\text{LiH}$  f.  $\text{Fe}_3\text{O}_4$  Hint; Identify the species being oxidized and reduced in each of the following reactions: a.

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